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USEFUL EQUATIONS molarity = $\frac{\text{mol solute}}{\text{L solution}}$

1 L = 1000 mL

The molarity of a solution is a ratio of the moles of

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- including limiting reagents and

percent yield Stoichiometry Basic

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214.2g OsF₃ x 1 mol OsF₃ = 12.9 M

OsF₃. 0.0673 L soln 247.23 g OsF₃.

Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution. 1.54 mol K₂SO₄ = 12.3 M K₂SO₄....

Molarity Worksheet 2 ANSWERS -

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The molar concentration can also be expressed as the following: 1.00 M

NaCl = 1.00 mol NaCl / 1 L NaCl

solution. and. 1.50 M Pb (NO₃)₂ =

1.50 mol Pb (NO₃)₂ / 1 L Pb (NO₃)₂

solution. First, we must examine the reaction stoichiometry in the balanced

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= 1 L solution mol solute. 1 L = 1000

mL. The molarity of a solution is a ratio of the moles of solute per liters of solution. The units for molarity are written as mol/L or M.

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seconds 86,546 views PRACTICE

PROBLEM: A 34.53 mL sample of
 H_2SO_4 reacts with 27.86 mL of
0.08964 M NaOH , solution , .

Calculate the , molarity , of Molarity
Dilution Problems Solution

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1 Worksheet #8 Solution

Stoichiometry and Chemical Reactions

Chem 102 Part I. Solution

Stoichiometry: Molarity, Dilutions and

Working with Density 1. What volume

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of a 6.00 M NaNO_3 solution should be used to make 0.525 L of a 1.20 M NaNO_3 solution. 2. To what volume (in L) should you dilute 0.200 L of a 15.0 M NaOH solution to obtain a 3.00 M NaOH solution? 3. An aqueous solution is prepared ...

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