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 Assuming the pressure stays constant, at what temperature will the volume of the gas be 604 cm³? Given: V1 = 506 cm³. T1 = 147 °C. V2 = 604 cm³. T2 = ? V1/T1 = V2/T2. 506 cm³/420 K = 604 cm³/T2. T2 = 501 K . or. 501 - 273 = 228 °C. Remember to convert temperature to Kelvin (420 K) Episode 902