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~~Stoichiometry: Limiting~~
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Problems and Examples

~~9.2 Ideal Stoichiometric
Calculations~~

~~Stoichiometry~~

~~Limiting \u0026amp; Excess~~

~~Reactant, Theoretical~~

~~\u0026amp; Percent Yield~~

Chemistry

The Periodic Table:

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~~& Molecular~~

~~Formula Problems~~

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*to Oxidation Reduction
(Redox) Reactions*

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Worksheet Answers~~

Chapter 9 Worksheet
Stoichiometry 1.

Consider the following
equation: $\text{As}_2\text{O}_3 + 2 \text{Cl}_2 + 5 \text{H}_2\text{O} \rightarrow 2 \text{H}_3\text{AsO}_4 + 4 \text{HCl}$
a. How many
moles of H_3AsO_4 can be
produced from 3.37 mol
of As_2O_3 ?

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~~Solved: Chapter 9~~

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~~Stoichiometry 1.~~

~~Consider The ...~~

Stoichiometry

Worksheet Answers

CHAPTER 9 REVIEW

Stoichiometry

SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1.

_____ The actual yield

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of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are

~~Chapter 9 Review~~

~~Stoichiometry~~

~~Worksheet Answers~~

Reaction stoichiometry is the subject of this chapter and it is based on chemical equations and the law of

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conservation of

matter. All reaction-
stoichiometry

calculations start with a
balanced chemical
equation. This equation
gives the relative
numbers of moles of
reactants and products.

~~CHAPTER 9~~

~~Stoichiometry—Quia~~

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Answers Chapter 9

Review Stoichiometry

Worksheet Answers

Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

398-399 #1-6, 9, 12 and answers Lesson 7:

Wednesday 22nd

January Unit 5 Test Unit

5 Binder Check. 3

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continued The
decomposition of N_2O_5
proceeds according to
the following equation:.

~~Unit 9 Stoichiometry
Worksheet Answers~~

Chapter 9: Standard
Review Worksheet 1.

Answers will vary. An
example is included

below: $2\text{H}_2\text{O}_2(\text{aq})$
 $2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This
describes the

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decomposition reaction
of hydrogen peroxide.

Microscopic: Two
molecules of hydrogen
peroxide (in aqueous
solution) decompose to
produce two molecules
of liquid water and one
molecule of oxygen gas.

~~Chapter 9: Standard~~

~~Review Worksheet~~

Chapter 9 Review

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~~Chapter 9 Review~~

~~Stoichiometry~~

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Worksheet Answers

- Stoichiometry is based on the law of conservation of mass.
- The mass of reactants equals the mass of the products.

~~Chapter 9~~

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SECTION 3

Worksheet
Answers

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H

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5.0 g Cu 1 mol Cu 1
mol Ag 107.9 g Ag =
8.5 g Ag. 63.5 g Cu 1
mol Cu 1 mol Cu. 8.5 x
100 = 55.9 % yield. 15.2

CHAPTER 11:
STOICHIOMETRY.

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**MOLE TO MOLE
RATIO.** When nitrogen
and hydrogen gas are
heated under the correct
conditions, ammonia
gas (NH_3) is formed. a.
RXN: $1. \text{N}_2 + 3. \text{H}_2 \rightarrow 2. \text{NH}_3$. b.

~~CHAPTER 11:~~ ~~STOICHIOMETRY~~

Answer Key. Chapter 9:
Standard Review
Worksheet. 1. Answers

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will vary. An example is ... We determine the theoretical yield by stoichiometric

calculations. The actual yield ... All Wkst KEYS and Review KEY Ch 9.pdf. Ch 9 – Worksheet 1 Using Stoichiometry.

From the equation :



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Chapter 9 Section 1

Intro to Page 2 ...

Chapter 9 Stoichiometry

Section 1 Answers

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